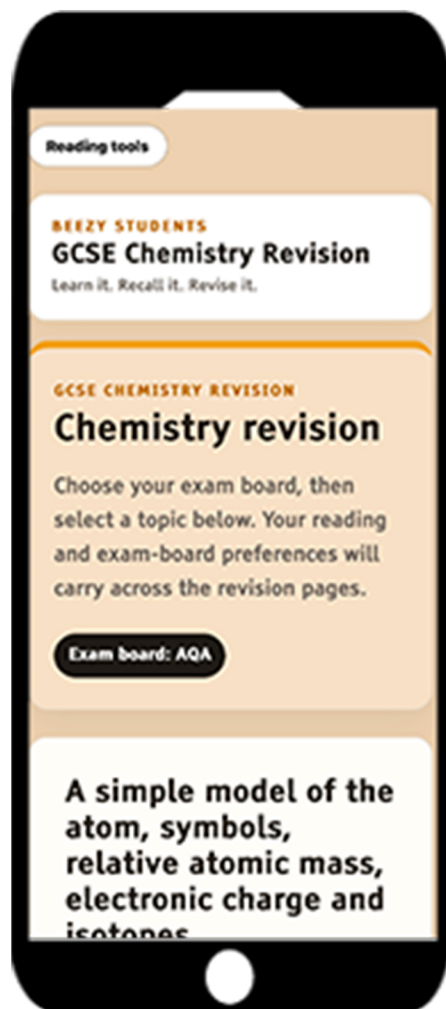
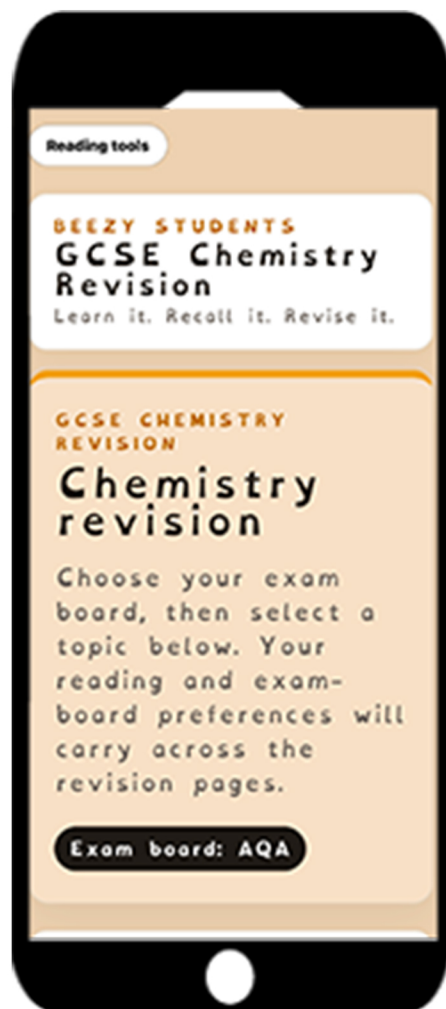
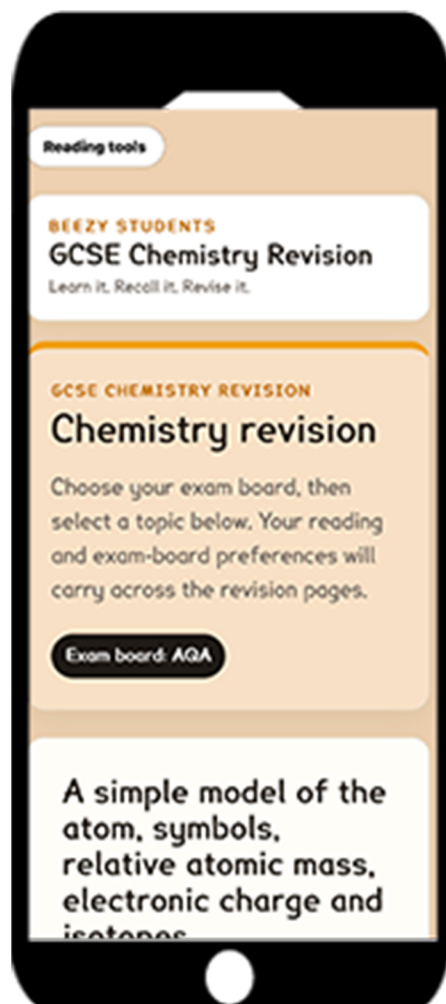
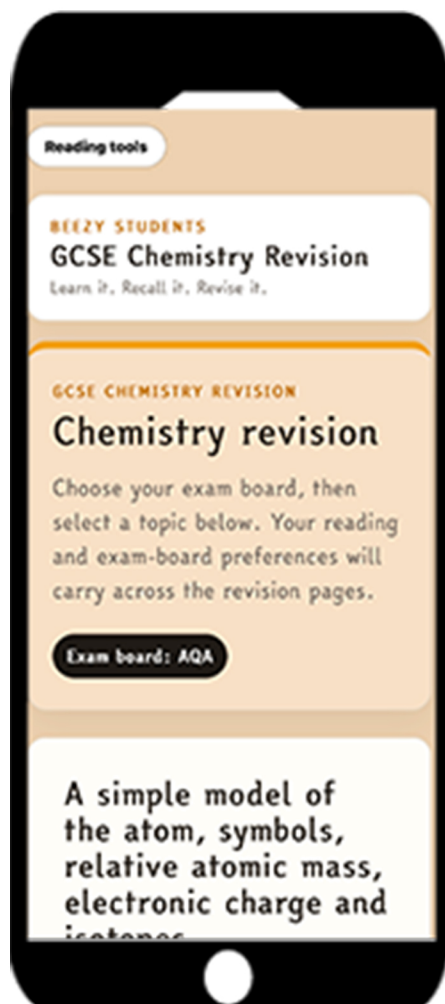
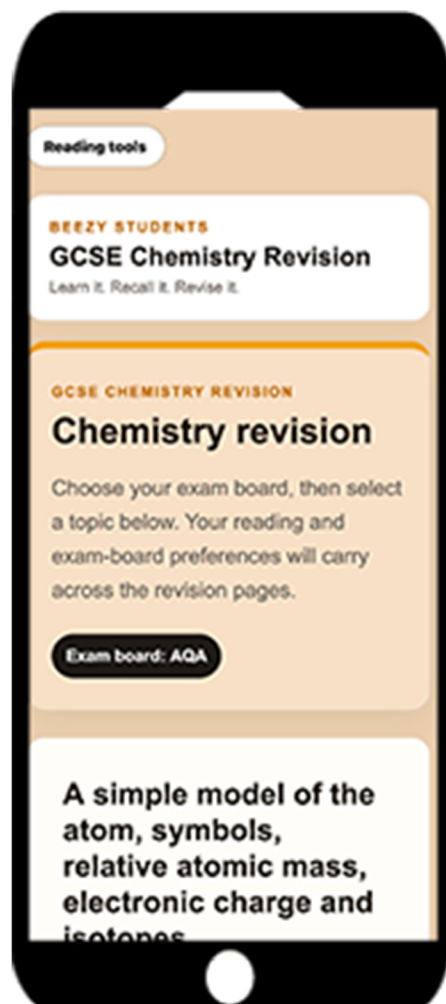




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1. Structure of an Atom

(AQA - AQA)

MAIN POINTS / CUES

NOTES

MY SUMMARY

2. The Periodic Table

(AQA - AQA)

MAIN POINTS / CUES

NOTES

MY SUMMARY

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4 sheets of paper

Destination

Save to PDF

Pages

All

Colour mode

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Pages per sheet

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Margins

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3. Atomic Number

(AQA - AQA)

MAIN POINTS / CUES

NOTES

MY SUMMARY

4. Subatomic Particles

(AQA - AQA)

MAIN POINTS / CUES

NOTES

MY SUMMARY

5. What the Subatomic Particles

(AQA - AQA)

MAIN POINTS / CUES

NOTES

MY SUMMARY

6. Subatomic Particles

(AQA - AQA)

MAIN POINTS / CUES

NOTES

MY SUMMARY

Print

1 sheet of paper

Destination

Save to PDF

Pages

All

Colour mode

Colour

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BEEZY STUDENTS

GCSE Chemistry Revision

Learn it. Recall it. Revise it.



GCSE CHEMISTRY REVISION

Atoms, elements and compounds

A simple model of the atom, symbols, relative atomic mass, electronic charge and isotopes

Cambridge 3.2.2 | 3.3.3 (E) | 2.2.6 | 8.1.1 | 8.1.4

Includes HT-only content

YOUR SPECIFICATION

Cambridge student objectives

- 3.2.2 — Define relative molecular mass, M_r , as the sum of the relative atomic masses. Relative formula mass, M_r , will be used for ionic compounds
- 3.3.3 (E) — Use the relationship $\text{mass (g)} \div \text{amount of substance (mol)} = \text{molar mass (g / mol)}$ to calculate: (a) amount of substance (b) mass (c) molar mass (d) relative atomic mass or relative molecular / formula mass (e) number of particles, using the value of the Avogadro constant
- 2.2.6 — State that: (a) Group VIII noble gases have a full outer